

Announcements

- Turn on the Clicker (the red LED comes on).
- Push “Join” button followed by “20” followed by the “Send” button (switches to flashing green LED if successful).
- Official exam 3 scores will be posted by noon today. (e-mail notification)
- Average was the same as first two exams.
- Don't forget to download Electrochem laboratory assignment parts. **You must turn in the completed prelab at the beginning of class.**
- Please check your quiz to see if you had circled "towards products" on question 2d. If you did and it is marked wrong please return it to me for regrading.
- **HAVE A GREAT THANKSGIVING BREAK!**

Electrochemistry

(Chpt 19 & sec 4.4)

- Voltaic cells
- Redox RXNs (and balancing)
- Energetics of Redox ($\Delta G = -nFE$)
- Standard half-cell potentials (E_o)
- Concentration and voltaic cell discharge
- Energy capacity (power) of batteries
- Electrolysis (recharging batteries)
- Electrochem and low emission vehicles.

Rules for assigning Oxidation #s

- All pure elements have an oxidation number = 0.
- O atoms in compounds usually have an oxidation number of -2, except in the case of peroxides.
- Alkali metals: +1, Alkali earths: +2, Halogens: -1 (but in oxides +1, +2....., in ClO^- , Cl is +1)
- H atoms in compounds have oxid# = +1, except in metal hydrides.
- The total of all the charges (oxid #s) on all the atoms in a molecule or ion must add up to the total charge on the species.
- Do not confuse oxid# with formal charge which is used to find the best Lewis structure.

Have a

Great

Thanksgiving Break