

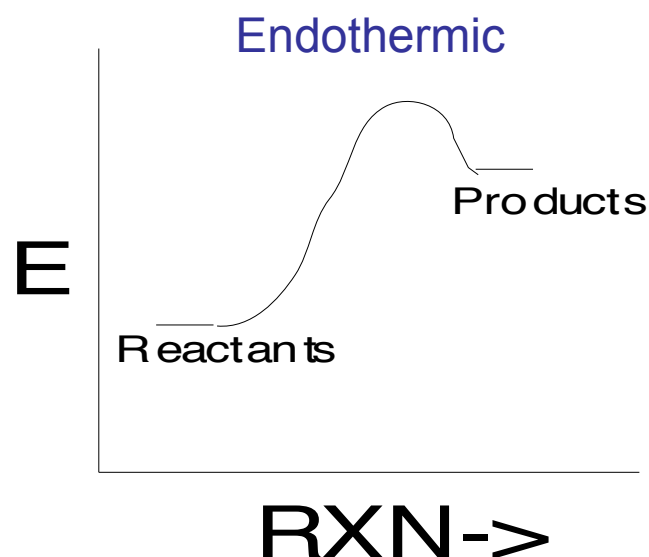
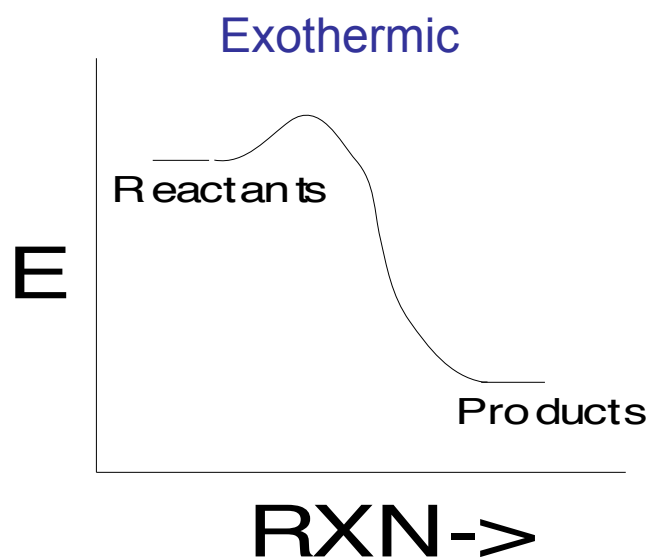
Announcements

To join clicker to class today:

- Turn on the Clicker (the red LED comes on).
- Push “Join” button followed by “20” followed by the “Send” button (switches to flashing green LED if successful).
- Syllabus and first reading assignment available on class web site.
- Class Password: .
- Deadline in syllabus for subscription to e-mail list was erroneous--should be 1 week from Friday(Sept. 15). This has been updated on the syllabus now available.
- If you are in the 1:50 discussion and could switch to 3 P please let me know. 1:50 discussion is overloaded.
- Wear appropriate clothes to Lab (No open toed shoes, skirts, shorts, etc.)

Review

- Thermochemistry = study of energy in chemical reactions.
- Exothermic processes release heat (lower chemical potential energy of the system)
- Endothermic processes absorb heat (raise the chemical potential energy of the system)
- Rates (kinetics) will be determined by what happens in the middle of a reaction process. (RXN path energy diagrams).



Naming Alkanes

- Special names for 1-4 carbons.
- 5 or more named by using the appropriate greek prefix + -ane.
- Note Bp increases with > # C.

TABLE 11.1 Alkanes in Natural Gas

Compound	Typical Abundance (% by Volume)	Formula	Lewis Structure	Boiling Point (°C)
Methane	75-90	CH ₄	<pre> H H-C-H H</pre>	-164
Ethane	5-15	C ₂ H ₆	<pre> H H H-C-C-H H H</pre>	-89
Propane	2-5	C ₃ H ₈	<pre> H H H H-C-C-C-H H H H</pre>	-42
Butane	<3	C ₄ H ₁₀	<pre> H H H H H-C-C-C-C-H H H H H</pre>	0